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Synthesis and Magnetic Behaviour of $p\text{-O}_2\text{NC}_6\text{F}_4\text{CNSSN}$

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The structure and magnetic properties of the dithiadiazolyl (DTDA) radical $p\text{-O}_2\text{NC}_6\text{F}_4\text{CNSSN}$, **1**, is described.

Keywords: dithiadiazolyl; ferromagnet

INTRODUCTION

We have been interested for some time in studying DTDA radicals as molecular magnetic materials. One of the long term goals of metal-free (organic) magnets is the design of materials which may have specialist uses in the electronics industry.

Unfortunately, most DTDA's exist as spin-paired dimers in the solid state, rendering them diamagnetic. The dimerisation energy is considerable (ca. 35 kJ mol⁻¹) and only two monomeric DTDA derivatives have been reported; $p\text{-NCC}_6\text{F}_4\text{CNSSN}$ ¹ and $p\text{-BrC}_6\text{F}_4\text{CNSSN}$. ²

Here we report the design and characterisation of $p\text{-O}_2\text{NC}_6\text{F}_4\text{CNSSN}$, **1**. ³

DESIGN OF 1

There exists a strong correlation between the Hammett parameter (σ_p) of the *para* substituent of the aromatic ring and the redox behaviour of the radical.⁴

Previously reported monomeric DTDA derivatives of formula $p\text{-XC}_6\text{F}_4\text{CNSSN}$ ($X = \text{CN}, \text{Br}$) have high σ_p values, while those that form dimers ($X = \text{Me}_2\text{N}, \text{MeO}, \text{F}, \text{Cl}$) have low or negative σ_p values. It was on the basis of this that the nitro derivative, **1**, was chosen as a molecule which would remain monomeric in the solid state.

STRUCTURE OF 1

$p\text{-O}_2\text{NC}_6\text{F}_4\text{CNSSN}$, **1**, may be prepared from $p\text{-O}_2\text{NC}_6\text{F}_4\text{CN}$ following standard synthetic techniques.⁴

As stated earlier, **1**, is monomeric in the solid state with molecules linked into chains *via* electrostatic $\text{NO}_2^{\delta-} \dots \text{S}^{\delta+}$ interactions (Fig. 1). There are no close $\text{S} \dots \text{N}$ interactions between heterocyclic rings.

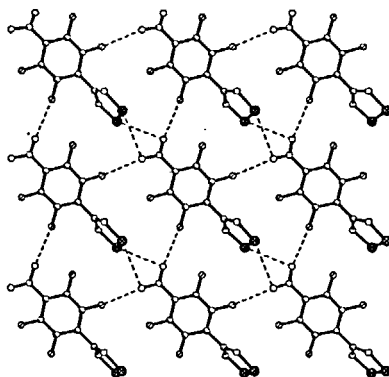


FIGURE 1: Crystal structure of **1** viewed down z axis

MAGNETIC PROPERTIES OF 1

1 is found to be paramagnetic in the solid state and a plot of inverse susceptibility against temperature (Fig. 2) indicates that it exhibits

Curie-Weiss behaviour with a Weiss constant of +1.59 K, indicating local ferromagnetic exchange.

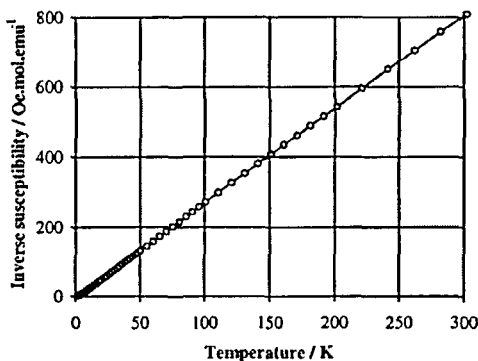


FIGURE 2: Inverse susceptibility plot for 3

The Curie-Weiss law fails at low temperature, possibly due to short range interactions among magnetic moments.

The effective magnetic moment of **1** is found to be $1.73\mu_B$ (except at low temperature), consistent with an $S = \frac{1}{2}$ paramagnet. The slight increase in magnetic moment at low temperature is due to the ferromagnetic interactions between radical centres. On cooling below 5.5 K the magnetic moment drops rapidly indicative of antiferromagnetic exchange interactions occurring between ferromagnetic sublattices.

DISCUSSION

The solid state structure of **1** suggests that the magnetic behaviour may be low dimensional. Preliminary studies find that the fit to a one-dimensional Heisenberg system requires an antiferromagnetic inter-chain term of similar magnitude to the intra-chain term, indicating **1** is not a good one-dimensional system. A better fit may be obtained using a two-dimensional Heisenberg model with a single exchange parameter, $J = 1.2$ K.

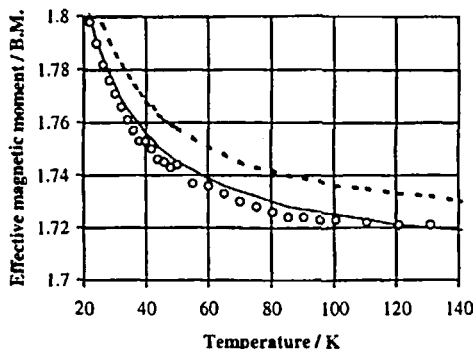


FIGURE 3: Modelling of magnetic behaviour of **1**
 -- = 1-D model, — = 2-D model, o = magnetic data

CONCLUSION

The first rationally designed monomeric DTDA radical is reported. Preliminary magnetic measurements indicate local ferromagnetic interactions between molecules. Further studies are planned.

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References

- [1] A. J. Banister, N. Bricklebank, I. Lavender, J. M. Rawson, C. I. Gregory, B. K. Tanner, W. Clegg, M. R. J. Elsegood and F. Palacio, *Angew. Chem. Int. Ed. Engl.*, 1996, **35**, 2533.
- [2] G. Antorrena, J. E. Davies, M. Hartley, F. Palacio, J. M. Rawson, J. N. B. Smith and A. Steiner, *J. Chem. Soc. Chem. Commun.*, 1999, 1393.
- [3] R. J. Less, C. M. Pask, J. M. Rawson, F. Palacio, P. Olliet and N. Feeder, *Manuscript in preparation*.
- [4] C. M. Aherne, A. J. Banister, I. B. Gorrell, M. I. Hansford, Z. V. Hauptman, A. W. Luke and J. M. Rawson, *J. Chem. Soc., Dalton Trans.*, 1993, 967.